Close the flask with a watch glass and allow to stand at laboratory temperature for 30 minutes. Add 20 ml. 10% Kl solution, shake and add 100 ml. of distilled water. Titrate with 0.1-N $Na_2S_2O_3$ solution using starch solution as indicator.

Repeat the test using the same quantities of reagents but allowing 3 hours for the absorption.

Carry out blank tests without oil alongside each determination.

Calculate in the usual manner:—
I.V. $_{30\ min.}$ = the Wijs iodine value obtained after 30 minutes' absorption.

I.V. $_{3\ hours}$ = the Wijs iodine value obtained after 3 hours' absorption.

Then the "instantaneous" iodine value is given by the expression:
I.V. $_{30\ min.}$ — 1/5 (I.V. $_{3\ hours}$ — I.V. $_{30\ min.}$)

We may add that closure of flasks by watch glasses during iodine value determinations has been practised in this laboratory for over 25 years with completely satisfactory results. If any halogen is lost during the period of standing we have yet to obtain proof that it can in anyway affect the results, while the procedure avoids irregularities which may arise from the presence of an aqueous solution in close proximity to the reagents.

In our experience iodine values so determined agree with those carried out with the greatest care in stoppered flasks.
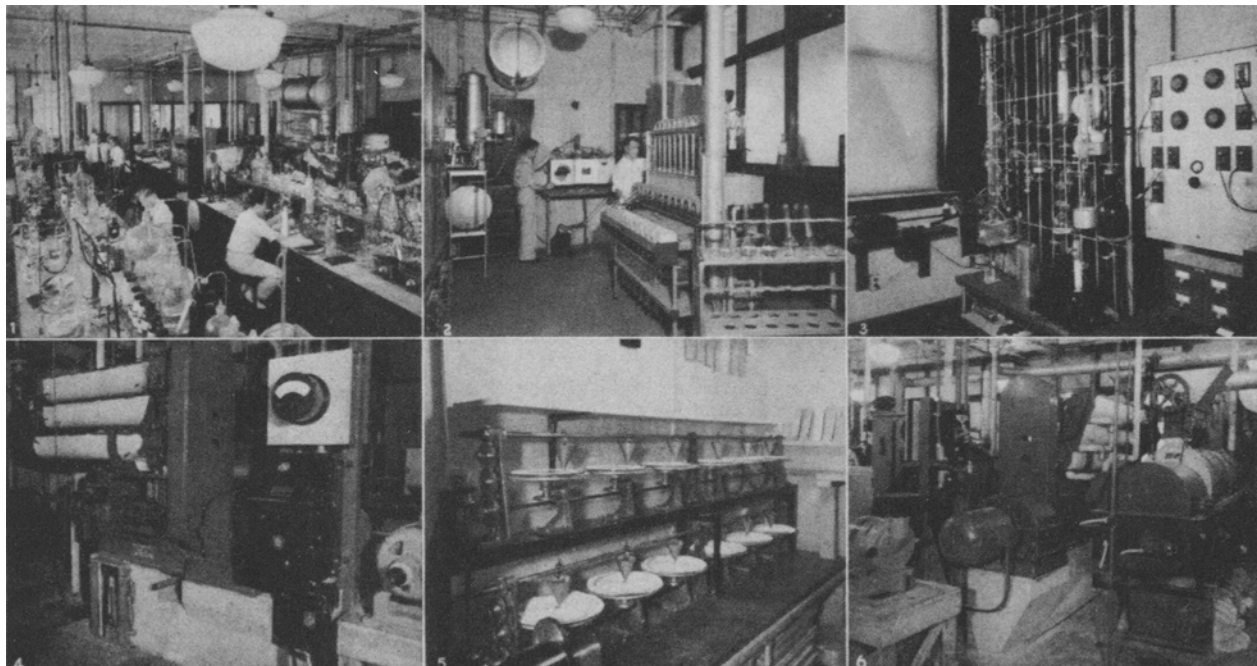
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1. Research on soybean meal and oil
2. Analytical equipment
3. High vacuum molecular stills and vacuum pumps
4. Small expeller, Engineering Section
5. Constant temperature-humidity room showing Sanderson drying meter and film spinner used in paint and varnish studies
6. Rotary drier, expeller and laboratory mill